

NAMING IONIC AND COVALENT COMPOUNDS

Syllabus reference 8.2.3

1 Name the following compounds:

- a CaO _____
- b AlCl_3 _____
- c Ag_2O _____
- d K_2SO_4 _____
- e NH_4Br _____
- f $\text{Mg}(\text{OH})_2$ _____
- g $\text{Ba}(\text{NO}_3)_2$ _____
- h Na_2O _____
- i FeCO_3 _____
- j PbS_2 _____

2 Write a formula for each of the following compounds:

- a magnesium chloride _____
- b calcium carbonate _____
- c copper(II) nitrate _____
- d sodium hydroxide _____
- e ammonium chloride _____
- f potassium sulfide _____
- g iron(II) oxide _____
- h lead(IV) bromide _____
- i silver nitrate _____
- j aluminium oxide _____

3 Name the following compounds:

- a PCl_3 _____
- b SF_4 _____
- c N_2O_3 _____
- d NO _____
- e Cl_2O_7 _____
- f HCl _____
- g CO_2 _____

4 Write the formula for each of the following compounds:

- a boron trichloride _____
- b phosphorus pentaiodide _____
- c hydrogen bromide _____
- d dinitrogen tetroxide _____
- e silicon dioxide _____